

## METHOD FOR THE SEPARATION OF ZEOLITES

The present invention relates to a method for  
10 processing synthetic zeolites, i.e. for recovering  
zeolite crystals, also having dimensions of less than  
500 Å, from solid/liquid dispersions containing them.

Zeolites can be used as molecular sieves and as  
catalysts in refining and in the chemical and  
15 petrochemical industry. It is known that the catalytic  
performances of zeolites are correlated to their  
crystalline structure and their composition. With the  
same parameters, however, very different performances  
can be obtained, depending on the morphology of the  
20 crystalline phase. The activity of a zeolitic catalyst  
increases with a decrease in the dimensions of the  
crystals: this probably depends on diffusive factors  
which influence the reactivity or favour the formation  
of heavy products which cause fouling, i.e. obstruction  
25 of the zeolitic pores. GB 1402981, for example,

describes the increase in the stability of ZSM-5, characterized by crystallites having dimensions of 0.005-0.1  $\mu\text{m}$ , in the transformation of hydrocarbons.

A.J.P.H. van der Pol et al., in Appl. Catal. A 92  
5 (1992) 113, correlate the crystal dimension of zeolites with the catalytic activity, in the hydroxylation of phenol with TS-1. EP 242960 describes that by using metal-silicates with a zeolitic structure, characterized by an external surface area  $> 5 \text{ m}^2/\text{g}$  and  
10 crystallites  $< 0.5 \mu\text{m}$ , the yields to  $\epsilon$ -caprolactam in Beckmann's catalytic rearrangement are improved.

The formation of crystals with varying dimensions can be obtained by suitably modulating the composition of the synthesis gel and hydrothermal treatment  
15 conditions (F. Di Renzo, Catalysis Today, 41 (1998) 37). The production of zeolites with small dimensional crystals is generally obtained with synthetic methods which tend to favour the nucleation process with respect to that of the growth of crystals.

20 The zeolitic phase, however, cannot be used as such as catalyst. If, on the one hand, the small dimensions of the zeolite crystallites favour the intraparticle diffusion of the reagents and reaction products and allow good catalytic performances to be  
25 obtained, on the other hand, similar dimensions are not

compatible with normal industrial reactor solutions. To overcome this problem, the zeolites are bound with suitable ligands, selected from oxides and their mixtures (for example alumina, silica-alumina, silica) or clays. The preparation methods of bound zeolites must be such that they do not cause blockage of the zeolitic cavities which would obviously result in a decrease in the catalytic activity.

The possibility of simplifying the overall preparation of the catalyst made up of the active phase and ligand is described for example in EP 906784 and MI 99 0024538 for the preparation of catalysts in spheres for applications in slurry or fluid bed reactors. In the processes described, the separation phase of the zeolite is avoided.

In other cases, above all when further treatment on the zeolitic phase in powder form is necessary, as, for example, when the zeolite must be used in extruded form in a fixed bed reactor, the separation of the zeolitic phase from the mother liquor or washing water or possible ionic exchange water during the processing or preparation of a catalyst, cannot be avoided.

The separation of the zeolite is, from an industrial point of view, a problem, especially when operating with a "small" crystal zeolite (submicronic

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agglomerates, mainly from 500 to 5000 Å; crystals from 100-500 Å), which cannot be separated from the synthesis medium with the usual techniques. There are numerous main parameters which can make the separation of the crystalline phase critical. The following can be mentioned for example: the intrinsic dimension of the crystals and/or agglomerates, their dimensional homogeneity, the presence of silica in the colloidal state in the mother liquor, as a result of a crystallization yield of less than 100% with respect to SiO<sub>2</sub>.

Zeolites with large crystals can be separated from a liquid phase by filtration in reasonable times (for example the declared filtering threshold for pressure filters, which depends on the cloths used, can vary from 5 to 300 µm), whereas this is not always possible for a zeolite with small crystals, as either the solid agglomerates on the filter, excessively slowing down the passage of the liquid through the panel formed thereby or it passes through the filter without being withheld.

In these cases, the solid-liquid separation can be obtained by means of centrifugation, but this method is long and costly. In addition, although the solid panel separated from the centrifuge is apparently compact, it

gives marked thixotropic phenomena. This means that the solid, when handled, takes on the characteristics of a liquid, with all the drawbacks relating thereto.

There are also methods based on micro-filtration which use membrane cartridges, such as those described for example in patent WO93/06917. In all cases, these methods require the use of specific equipment for the purpose.

A simple, rapid and inexpensive process has now been found for effecting solid-liquid separations also in the case of small crystal zeolites. With this new method, the zeolite crystals, also with dimensions of less than 500 Å, can be isolated by means of simple decanting or filtration which can be carried out in very fast times.

The object of the present invention therefore relates to a process for separating zeolite crystals from aqueous environments containing them in suspension, which comprises:

- (a) treating this suspension of crystals with an acid up to a pH ranging from 3 to 8;
- (b) subjecting the resulting mixture to filtration or decanting to isolate the crystals.

In particular, this process can be applied to zeolite crystals in suspension in the crystallization

mother liquor and, by modification of the composition of the mixture, it allows the zeolitic phase to be separated from the mother liquor by filtration or decanting.

5       A particular aspect of the present invention is therefore a process for the recovery of zeolite crystals in suspension in the crystallization water which comprises:

- 10       (a) treating this suspension of crystals with an acid up to a pH ranging from 3 to 8;
- (b) subjecting the resulting mixture to filtration or decanting to isolate the crystals.

      For the process of the present invention, organic or inorganic, Broensted or Lewis type acids can be used. Suitable acids which can be used are acetic acid, hydrochloric acid, nitric acid, formic acid, propionic acid and oxalic acid. The pH of step (a) preferably ranges from 3 to 6. The acids are used in aqueous solutions having a concentration ranging from 0.05 to 10 N.

      All zeolites are suitable for being separated from the crystallization mother liquor according to the method of the present invention. For example, zeolites with small pores which can be separated according to the method of the present invention are ANA, CHA, ERI,

GIS, LEV, LTA, MTN, PHI, RHO, SOD; zeolites with medium pores can be EUO, FER, MFI, MEL, MTT, TON, MFS, NES; zeolites with large pores can be BEA, FAU, MTW, MOR, OFF, MAZ, LTL, GME, EMT; zeolites with extra-large pores can be CFI and DON.

Another particular aspect of the present invention is that the crystalline phase is separated in a mixture with oxides which can be used as ligands in the subsequent extrusion treatment.

Said oxides can be generated by the reagents of the zeolite preparation reagent mixture, not transformed into crystalline phase during the synthesis hydrothermal treatment, or they can be added to the suspension of crystals before the separation treatment, or again they can be generated by precursors suitably added to the crystallization slurry.

In the first case, the method has the additional advantage of quantitatively exploiting the silica, alumina and/or other metal oxide sources used in the preparation phase of the reagent mixture.

In the second case, an aqueous suspension containing clay, a silicon oxide, a silica-alumina or an alumina (for example bohemite) is also added to the zeolite suspension in the mother liquor.

In the third case, precursors of silica, silica-

alumina, alumina, such as for example aluminum acetylacetonate, alkylaluminates and/or alkylsilicates are added to the suspension of zeolite crystals in the mother liquor, which, by addition of the acid according to the present invention, produce a rapid aggregation of the colloidal phases formed and the precipitation of the corresponding oxides. According to a particular aspect of the present invention, the acid in step (a) can be added by means of a precursor capable of contemporaneously generating, by hydrolysis, both the acid and an oxide suitable for being used as a ligand. Appropriate precursors for this purpose can be  $\text{Al}(\text{NO}_3)_3$ ,  $\text{Al}(\text{SO}_4)_3$ , silicic acid, silicon or aluminum halides,  $\text{Al}(\text{CH}_3\text{COO})_3$ .

15       The quantity of oxides co-present with the zeolite, which are isolated at the end of the separation process according to the particular aspects described above of the present invention, can vary from 1 to 50% by weight with respect to the zeolite.

20       After the treatment with acid up to a pH ranging from 3 to 8, and optional addition of ligand oxides, or their precursors, demineralized water can be added to the mixture, which is stirred, left to deposit and the supernatant liquid separated. In this way, it is possible to effect solid-liquid separations by

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decanting, using a simple method and without having to resort to particular equipment.

According to another possibility, the same separations can be carried out by filtration instead of  
5 by decanting.

A further advantage of the method described consists in the fact that the acid treatment also allows a partial exchange of the zeolite into acid form, in a single passage. In this phase, a zeolite  
10 can, in fact, be obtained in partially acid form, leaving a contact time which is sufficient to effect the ionic exchange between the  $H^+$  ion and the alkaline or earth-alkaline metal present in the zeolite. The entity of the exchange depends on the accessibility of  
15 the cationic exchange sites; for example for zeolites with large pores, it is at least 30%.

This allows a significant reduction in the unitary operations required for the processing of zeolites, processing referring to the washing of a zeolite to  
20 purify it from the crystallization mother liquor and all the subsequent operations for obtaining catalysts, such as ionic exchanges, for example.

When operating according to the known art, the processing of a zeolite to obtain its acid or ammonia  
25 form comprises the following operations:

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1. Separation of the solid crystalline phase from the crystallization slurry by means of filtration or centrifugation;
  2. Washing by re-dispersion in water;
  - 5 3. Separation of the zeolite by means of filtration or centrifugation;
  4. Drying;
  5. Calcination to remove the organic templating agent (e.g. quaternary alkylammonium ion);
  - 10 6. Exchange in aqueous solution containing an acid or an ammonium salt;
  7. Separation of the zeolite by means of filtration or centrifugation, accompanied by washing to remove the excess acid or ammonium salt and the corresponding  
15 exchange products;
  8. Drying;
  9. Calcination to remove the ammonium ion, when the exchange has been carried out with an ammonium salt and the acid form is required.
  - 20 Operations 6 and 7 are normally repeated in relation to the exchange level desired. For residual alkaline metal values lower than 150 ppm they are normally repeated at least twice.

Operating according to the method of the present  
25 invention, the sequence of processing operations for

obtaining the zeolite in ammonia or acid form becomes:

(a) treating the suspension of zeolite crystals in the crystallization mother liquor containing them with an aqueous solution of acid up to a pH ranging from 3 to

5 8;

(b) subjecting the resulting mixture to decanting or filtration to separate the zeolite crystals;

(c) drying;

(d) calcination to remove the templating agent;

10 (e) exchange in an aqueous solution containing an acid or ammonium salt;

(f) separating the zeolite by means of filtration or decanting, accompanied by washing to remove the excess acid or ammonium salt and the corresponding exchange

15 products;

(g) drying;

(h) calcination to remove the ammonium ion, when the exchange has been carried out with an ammonium salt and the acid form is required.

20 The zeolite crystals separated under point (b) do not require further washing.

Thanks to the exchange contribution of phase (a), steps (e) and (f) must be effected only once to guarantee residual alkaline or earth-alkaline metal  
25 values lower than 150 ppm.

When operating according to the method of the present invention, there are fewer and more rapid separation operations as the acid treatment effected speeds up both the filtration of the zeolite crystals  
5 from the suspension in the mother liquor, and also the subsequent filtrations following washing or ion exchange treatment.

In the particular case in which the zeolites are characterized by tridimensional large pore or extra-  
10 large pore systems, such as BEA, FAU, MOR, for example, the processing scheme is even more simplified:

- (a) treating the suspension of zeolite crystals in the crystallization mother liquor containing them with an acid up to a pH ranging from 3 to 8;
- 15 (b) subjecting the resulting mixture to decanting or filtration to separate the zeolite crystals;
- (c) exchanging the zeolite with an aqueous solution containing an acid or ammonium salt (the calcination phase to remove the organic templating agent is not  
20 necessary);
- (d) separating the zeolite by filtration or decanting;
- (e) drying;
- (f) calcination to remove the residual templating agent and the ammonium ion, when the exchange has been  
25 carried out with an ammonium salt.

The zeolite crystals obtained according to one of the two schemes provided above can be subsequently subjected to extrusion, after the addition of a suitable ligand, to obtain catalysts which can be used in fixed bed reactors.

In all the cases considered, the last calcination phase can be effected before or after extrusion.

When, in step (a) of the above schemes, an oxide suitable for being used as a ligand, or a precursor of said oxide or an acid precursor and oxide, has been added, the resulting mixture of step (e) can be directly fed to the extruder even without the further addition of a ligand.

Also when, in step (a), the precipitation of the zeolite preparation reagents not transformed into crystalline phase, is obtained, the resulting mixture of step (e) can be directly fed to the extruder even without the further addition of a ligand.

#### EXAMPLE 1 (Preparation of beta zeolite)

4.4 kg of tetra-ethylammonium hydroxide at 40% by weight, in aqueous solution, are added to 2.4 kg of demineralized water. 0.24 kg of sodium aluminate at 56% by weight of  $\text{Al}_2\text{O}_3$  are subsequently added, with continual stirring until a limpid solution is obtained. 6.0 kg of Ludox HS 40 colloidal silica at 40% of  $\text{SiO}_2$

are added to this solution. A homogeneous suspension is thus obtained, which is charged into an AISI 316 steel autoclave equipped with an anchor stirrer. The gel is left to crystallize under hydrothermal conditions at 180°C for 30 hours. The autoclave is cooled and the suspension of zeolite crystals in mother liquor is discharged for the subsequent processing step.

#### EXAMPLE 2

1 kg of the suspension obtained from the previous example is treated with about 300 g of 3 N acetic acid. The addition of the acid is effected under stirring, by subdividing the acid into several aliquots and leaving a few minutes between the addition of the various portions. The pH proves to be about 5. At this stage, a rather dense suspension is obtained to which 7 litres of demineralized water are added. When the stirring is interrupted, it can be observed that the whole solid is deposited in a few minutes, leaving a limpid supernatant liquid, and can be easily separated by filtration.

The solid thus obtained is analyzed with the XRD technique, and proves to consist of beta zeolite.

#### EXAMPLE 3

The beta zeolite crystals isolated by filtration as described in example 2, are re-dispersed in a

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solution of demineralized water and ammonium acetate, as described in the known art. The beta zeolite is left, under stirring, in contact with one of these solutions, at room temperature, for about 3 hours and the whole mixture is then sent for filtration on a common flat filter with a filter cloth, under nitrogen pressure. The panel is washed with water.

The humid panel obtained consists of beta zeolite in ammonia form.

10 To obtain the zeolite in acid form, the solid is dried at 150°C and is then calcined at 550°C for 5 hours in air.

Chemical analysis on the calcined zeolite gave the following results: Al 2.68%; Na 100 ppm; molar ratio Al/Na 228.

#### EXAMPLE 4

20 The beta zeolite crystals isolated as described in example 2 are calcined before being exchanged with a solution of ammonium acetate, filtered, washed and re-calcined as described in example 3. In this case, chemical analysis effected on the final zeolitic sample in acid form gave the following results: Al 2.57%; Na 106 ppm; molar ratio Al/Na 206.

25 On comparing the results of examples 3 and 4, it can be seen that it is possible to exchange the sodium

with  $H^+$  or  $NH_4^+$  without first eliminating the templating agent by calcination.

EXAMPLE 5 - Comparative

1 Kg of zeolite suspension in the mother liquor  
5 obtained in example 1 is processed according to the traditional procedure, i.e. by means of solid-liquid separation effected by filtration.

The filtration proves to be 5 times slower than  
the filtration carried out in example 2 after acid  
10 treatment.

The zeolite is washed with demineralized water  
until the washing water has a pH of 9. At this point,  
it is dried at  $150^\circ C$  and calcined for 5 hours in air at  
 $550^\circ C$ . The solid is then subjected to double exchange  
15 with ammonium acetate in aqueous solution (according to the procedure of the known art), with separation by filtration and washing with water.

The sample is then dried in an oven at  $150^\circ C$  and  
calcined for 5 hours at  $550^\circ C$  in air. Beta zeolite in  
20 acid form is thus obtained. Elemental chemical analysis on the latter sample gave the following results: Al 2.51%; Na 98 ppm; molar ratio Al/Na 218.

On comparing the results obtained in examples 3, 4  
according to the invention and comparative example 5,  
25 in relation to the quantity of residual sodium, it can



be deduced that the quality of the samples obtained is equivalent, but the procedure of the present invention allows a simpler and more rapid operation, as, in fact, there are fewer passages and the filtration times are shorter.

EXAMPLE 6 (Preparation of ZSM-12 zeolite)

185 g of tetra-ethylammonium hydroxide at 40% by weight, in aqueous solution, are added to 100 g of demineralized water. 4 g of sodium aluminate at 56% by weight of  $\text{Al}_2\text{O}_3$  are subsequently added. The limp solution thus obtained is poured, under stirring, into 500 g of Ludox HS 40 colloidal silica. After brief stirring, a limp and homogeneous gel is obtained, which is poured into an AISI 316 steel autoclave equipped with an anchor stirrer. The gel is left to crystallize under hydrothermal conditions at 160°C for 60 hours. At this point, the autoclave is cooled. The suspension obtained is homogeneous with a lactescent appearance.

EXAMPLE 7 (comparative)

100 g of the suspension of example 6 are filtered. The solid discharged is washed by re-dispersion in water, re-centrifuged, dried and calcined at 550°C.

Upon X-ray diffraction analysis, the solid obtained proves to consist of pure ZSM-12. Upon TEM

analysis the zeolite proves to have spherical crystalline agglomerates with dimensions of 0.1-0.3  $\mu\text{m}$ .

Chemical analysis provides a molar ratio  $\text{SiO}_2/\text{Al}_2\text{O}_3 = 94$  and  $\text{Na}/\text{Al} = 0.79$ . The treatment parameters and results obtained are indicated in Table 1.

#### EXAMPLE 8

100 g of the crystallization suspension of example 6 are treated with acetic acid, as described in example 2, obtaining a pH of about 5, and is then subjected to filtration to separate the solid which is subsequently dried and calcined. The sample obtained is characterized by the following molar composition:  $\text{SiO}_2/\text{Al}_2\text{O}_3 = 97$  and  $\text{Na}/\text{Al} = 0.4$ , thus revealing the removal of about 50% of the sodium present as counter-ion in the zeolite. The treatment parameters and results obtained are indicated in Table 1.

#### EXAMPLE 9

Example 8 is repeated using aluminum nitrate which by hydrolysis contemporaneously causes the formation of aluminates and nitric acid.

The test is carried out using a quantity of  $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$  equal to 27 g. The treatment parameters and results obtained are indicated in Table 1.

#### EXAMPLE 10

Example 9 is repeated with a quantity of

Al(NO<sub>3</sub>)<sub>3</sub>·9H<sub>2</sub>O equal to 14 g.

Table 1 below summarizes the tests effected. The acid compounds added to the slurry are indicated, together with its final pH and the time necessary for its filtration. The time necessary for the filtration of the crystallization suspension obtained in example 6 and not subjected to acid treatment according to the present invention, is provided as a comparison.

TABLE 1

EXAMPLES	Acid	pH	Filtration time (min)	solid % *
7	-	12.1	30	100
8	CH <sub>3</sub> COOH	4.9	5	104
9	Al(NO <sub>3</sub> ) <sub>3</sub>	3.1	13	117
10	Al(NO <sub>3</sub> ) <sub>3</sub>	6.0	3	111

\* normalized with respect to the solid obtained by centrifugation

From the results indicated in the table, it is evident that the acid, or acid precursor, added to the suspension of zeolite crystals in the crystallization mother liquor, favours the solid-liquid separation and allows the recovery of the crystals by filtration in much faster times.

This method is also extremely compatible with a subsequent forming process: it allows, in fact, the non-reacted silica during the crystallization process of the zeolite (example 8) and the alumina deriving from the hydrolysis of the acid precursor (examples 9 and 10), to be exploited as ligands.

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